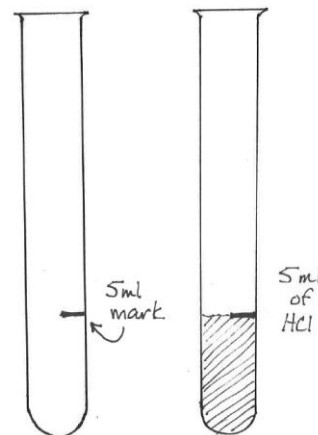


1. Use the graduated cylinder to measure 5 ml of water into one of the test tubes, and mark the level of the water with a pen. Dump out the water.
2. Get 5 ml of Hydrochloric Acid (HCl) solution from the teacher in the test tube, and place the test tube in the test tube rack.
3. Write the name and formula for each reactant in the table.

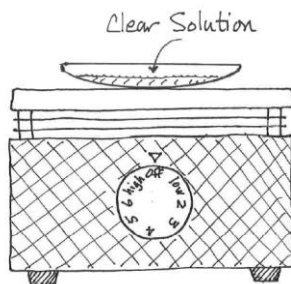
OBSERVE THE REACTION BETWEEN HYDROCHLORIC ACID AND MAGNESIUM.

4. Put the Magnesium strip into the solution, and notice the gas being produced.
5. Identify the gas by passing a flaming splint over the mouth of the test tube. Record the name and formula of the gas as one of the products.

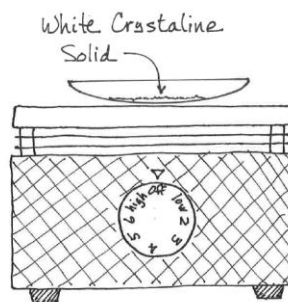


IDENTIFY THE SECOND REACTANT.

6. When the reaction is over, pour a small amount of the solution in the watch glass.
7. Put the watch glass on the hot plate, and let the water evaporate. There should be a white, crystalline solid on the watch glass.
8. Determine and record the name and formula of this product.
9. Clean everything up with water.



Before Water Evaporates



After Water Evaporates

SET UP FOR THE REACTION BETWEEN LEAD(II) CHLORIDE AND POTASSIUM IODIDE.

1. In one test tube, get a small amount of Lead (II) Chloride from the teacher.
2. In the other test tube, get a small amount of Potassium Iodide from the teacher.
3. Fill in the names, formulas and appearance of these two reactants in your data table.
4. Add a small amount of distilled water to each test tube, and record appearance of the reactants in water in your data table.
5. Without spilling the contents, set both test tubes in the 250 ml beaker.
6. Set the beaker on the triple-beam balance and record the mass of the beaker and the two test tubes of solution into **Mass Before Reaction**.

LEAVE THE BEAKER ON THE BALANCE.

OBSERVE THE REACTION AND MEASURE THE MASS OF THE PRODUCTS.

7. **Without moving the beaker**, pick up one test tube and carefully dump the contents into the other test tube.
8. Put the empty test tube back exactly where it was, and record the mass of everything now under **Mass After Reaction**.
9. Record the appearance of the solution now.
10. Allow the solution to stand overnight. The following day, pour a small amount of the solution in the watch glass and let the water evaporate. As in part 1, there should be a white, crystalline solid on the watch glass.

OBSERVE THE REACTION OF ALKASELTZER IN ROOM TEMPERATURE WATER.

1. Measure and record the mass of an AlkaSeltzer Tablet.
2. Put 200 ml of room temperature water into a 250 ml beaker and record its temperature.
3. Drop the AlkaSeltzer Tablet in the water and use the stopwatch to time how long it takes for the tablet to complete react (disappear). Record the time.
4. Dump the solution in the sink.

OBSERVE THE REACTION OF ALKASELTZER IN HOT WATER.

5. Measure and record the mass of a second AlkaSeltzer Tablet.
6. Put 200 ml of hot water from the teacher's station into a 250 ml beaker and record its temperature.
7. Drop the AlkaSeltzer Tablet in the water and use the stopwatch to time how long it takes for the tablet to complete react. Record the time.
8. Dump the solution in the sink.

OBSERVE THE REACTION OF ALKASELTZER IN ICE WATER.

9. Measure and record the mass of a third AlkaSeltzer Tablet.
10. Put 200 ml of ice water from the teacher's station into a 250 ml beaker and record its temperature.
11. Drop the AlkaSeltzer Tablet in the water and use the stopwatch to time how long it takes for the tablet to complete react. Record the time.
12. Dump the solution in the sink.

OBSERVE THE REACTION OF ALKASELTZER WHEN BROKEN INTO SMALL PIECES.

13. Measure and record the mass of a fourth AlkaSeltzer Tablet.
14. Put 200 ml of room temperature water from the teacher's station into a 250 ml beaker and record its temperature.
15. Break the AlkaSeltzer Tablet into 8-10 pieces.
16. Drop the pieces into the water and use the stopwatch to time how long it takes for the tablet to complete react. Record the time.
17. Dump the solution in the sink.