Integrated Science Objectives	Unit 7: Chemical Reactions

1. List the three characteristics of a chemical reaction.
2. Distinguish between a chemical change (reaction) and a physical change.
3. Define the terms Reactant and Product.
4. Name the Reactants and Products for the Photosynthesis reaction.
5. Write a flow chart for any chemical reaction.
6. Define the term Precipitate, and explain what a precipitate looks like in a chemical reaction.
7. Explain how you determine which product is the precipitate in a chemical reaction.
8. Explain why it's necessary to balance a chemical reaction in terms of the Law of Conservation of Mass.
9. Give an example of a chemical equation.
10. List the steps you go through when you write the chemical equation that describes a chemical reaction.
11. Distinguish between a subscript and a coefficient.
12. Explain what each of the following symbols mean: "s", "l", "g", "aq".
13. List the six Diatomic gases, and explain what a diatomic gas is.
14. Give an example of a Synthesis Reaction, and explain exactly what happens in this type of reaction.
15. Define the term "combustion".
16. Give an example of a Decomposition Reaction, and explain exactly what happens in this type of reaction.
17. Given any combination of Reactants; predict the products, write a flow chart for the reaction, and write the complete, balanced equation for that reaction.